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Please do not turn over this assessment until you are told to do so.

There are a total of 44 questions for a total of 100 points for exam 1. If there are multiple-choice questions, they will be marked, circled and the total will be converted to 100 points.

Questions 1-4 are from exam 11.

1. Determine the number of protons, neutrons, and electrons in a $^{235}\text{Pu}^{3+}$ ion.

- a. 94 protons, 141 neutrons, 91 electrons
- b. 94 protons, 141 neutrons, 94 electrons
- c. 94 protons, 141 neutrons, 94 electrons
- d. 94 protons, 141 neutrons, 94 electrons
- e. 94 protons, 141 neutrons, 94 electrons



2. A sample of a metal has a mass of 10.0 g.

- a. 10.0 g
- b. 10.0 g
- c. 10.0 g
- d. 10.0 g
- e. 10.0 g

3. According to the periodic table, the element with the highest atomic number is:

- a. Fr
- b. U
- c. Og
- d. Pu
- e. Hg

4. Which element is in the same group as oxygen?

- a. S
- b. Se
- c. Te
- d. Po
- e. At
- f. Bi
- g. Pb
- h. Tl
- i. Sn
- j. Sb
- k. Te
- l. I
- m. Xe
- n. Rn

5. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

6. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

7. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

8. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

9. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

10. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

11. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

12. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

13. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

14. Which of the following is not a noble gas? (a) He (b) Ne (c) Ar (d) Kr (e) Xe (f) Rn

Chapter 1 – Fundamentals of measurement and calculation

List the main metric prefixes and the meaning of each

For each number value given, underline the significant digits and list the correct number of significant digits

- .00342 _____
- 140000 _____
- 12.010 _____
- 28.0 _____

Convert each number given to proper scientific notation form

- .000048 _____
- 38700000 _____

Convert each number given into proper decimal form

- 6.84×10^{-4} _____
- 5.2×10^6 _____

Use dimensional analysis to convert between the metric units given:

- Convert 12.9m to cm
- Convert .025 kg to g
- Convert .00kg to mg

State Standard C.1.7

- o Define density and distinguish among materials based on densities. Perform calculations involving density.

A liquid has a mass of 32.1g and a volume of 43.8mL. Calculate the density.

Determine the mass (g) of a piece of wood with a density of .75 g/cm³ and a volume of 30.0 cm³

Determine the volume (mL) of a glass of water (density = 1.0 g/mL) that has a mass of 125g

CP CHEMISTRY HALF YEAR ASSESSMENT REVIEW TOPICS AND PROBLEMS

Answers will be posted on eboard
The back of each chapter has problems- by topic- work them!!!

Multiple Choice and Problems

STEPS OF SCIENTIFIC METHOD IN ORDER

FORMULA WRITING and NAMING- write a correct formula or name a compound correctly

- Na₂PO₄ HCl SO₂ Fe₂O₃
- Magnesium hydroxide carbon dioxide Copper (II) sulfate

GENERAL INFORMATION

Know the charges of atoms and common ions:

Symbol	charge
Ca atom	
Ca ion	
Sulfur	
Sulfide	

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

H	He						
Li	Be	B	C	N	O	F	Ne
Na	Mg	Al	Si	P	S	Cl	Ar
K	Ca	Ga	Ge	As	Se	Br	Kr
Rb	Sr	In	Sn	Sb	Te	I	Xe
Cs	Ba	Tl	Pb	Bi	Po	At	Rn
Fr	Ra						

1. Group _____ in the figure above contains only metals.
 - a. 2
 - b. 13
 - c. 17
 - d. 18
2. Based on their location in the figure above, oxygen and selenium have _____.
 - a. the same number of neutrons.
 - b. the same conductivity.
 - c. similar properties.
 - d. the same number of electron orbitals.
3. What is the atomic number for aluminum from the figure above?
 - a. 13
 - b. 14
 - c. 26.98
 - d. 26.9815
4. In the figure above, a neutral atom of silicon contains _____.

Chemistry Practice Exam #1

11.

Do the following conversions:

a) 54.2 cm^3 to in^3 (1 $\text{in} = 2.54 \text{ cm}$)	b) $1.794 \times 10^3 \text{ mm}$ to cm
c) 0.87 g/mL to lb/gallon (1 $\text{lb} = 453.6 \text{ g}$, 1 $\text{gallon} = 0.264 \text{ L}$)	d) $0.8547 \times 10^3 \text{ mg}$ to kg
e) $1.887 \times 10^3 \text{ L}$ to μL	

Chemistry 30B Exam 2 Practice Problems Solutions Page 1

- (a) In x-ray crystallography, x-ray patterns are measured or diffacted by regularly spaced atoms in a crystal.

(b) In an early but very important application of x-ray crystallography, the Bragg's measured (for the first time) carbon-carbon bond lengths and carbon-carbon bond angles in a substance called diamond.

(c) Rosalind Franklin verified that DNA is a helical molecule when she saw an X pattern in photo 51, one of many DNA diffraction patterns produced by x-ray crystallography.
- (a) No. X-ray diffraction locates an atoms space, and can give some idea as to whether or not it has a higher atomic number than other atoms due to the intensity of the scattering, but it cannot reveal what elements are present.

(b) Yes. X-ray diffraction reveals the relative position of atoms in space, so stereochemistry can be assessed.

(c) Yes. X-ray diffraction reveals the relative position of atoms in space, so distance between atoms anywhere in the molecule can be measured.

(d) Yes. X-ray diffraction reveals the relative position of atoms in space, so the relative position of three (or more) atoms can be evaluated.
- (a) Because ring planarity. Yes.

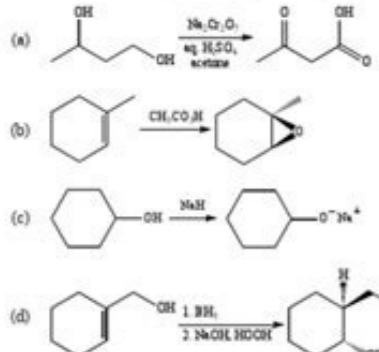
(b) Nitro group (-NO₂) resonance. No. X-ray diffraction may reveal that the NO₂ atoms lie in the same plane and that the N-O bond lengths are equal, but these facts alone are not proof of resonance. X-ray diffraction reveals the position of atoms, but not of electrons.

(c) Oxygen-nitrogen bond length. Yes.

(d) Presence of lone pairs on oxygen in nitro group. No.

(e) Electronegativity of oxygen > nitrogen. No.

(f) Positive formal charge on nitrogen in nitro group. No.



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The amount of energy required to raise the temperature of one gram of pure water by one degree Celsius is defined as: A calorie. 66 Show detail 1 day ago quizlet.com Show details a. the ability to carry an electric current well and to hold electric charge. b. taking up space and having mass. c. being brittle and hard. d. being malleable and ductile. An atom is. a. the smallest unit of matter that maintains its chemical identity. b. the smallest unit of a compound. c. always made of carbon. 410 Show detail 4 days ago quizlet.com Show details Chemistry Midterm Exam Review. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. chickenlittle. Terms in this set (95) Chapter 1. Matter and Change. Atom. it is the smallest unit of an element that maintains the chemical identity of that element. Element. 485 Show detail 1 week ago quizlet.com Show details Start studying Chemistry Midterm Exam review. Learn vocabulary, terms, and more with flashcards, games, and other study tools. 321 Show detail 1 week ago quizlet.com Show details Chemistry Midterm Exam Review. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. memanoear. all of that will be on the Midterm Exam. Terms in this set (32) What is the periodic law? 53 Show detail 1 day ago npsd.k12.nj.us Show details Chemistry - Mid Term Exam Review Sheet #1 The midterm exam covers chapters 1 - 4 & 9 - 11. You should read through each chapter, look over old tests you still have, answer the l-lowing questions and do the calculations in order prepare yourself lbr the mid-term. 1. Define the l-lowing terms and describe where each is located. proton - neutron - 294 Show detail 1 day ago njctl.org Show details Chemistry Midterm Review Questions Atoms, Molecules, Ions & Compounds 1) The nucleus of an atom contains _____. A) electrons B) protons, neutrons, and electrons C) protons and electrons E) protons 2) Potassium is a ____ and chlorine is a _____. 432 Show detail 1 week ago quizlet.com Show details 1.204×10^{23} atoms. A sample of tin (atomic mass 118.71 amu) contains 3.01×10^{23} atoms. The mass of the sample is. 59.3 g. The frequency of electromagnetic radiation is measured in waves/second, or. hertz. According to the particle model of light, certain kinds of light cannot eject electrons from metals because. 232 Show detail 1 week ago lcps.org Show details Chemistry Semester Exam Review. Complete all problems on notebook paper. Transfer your final answers to the answer sheet. Measurements and Calculations. ... Chemistry Midterm Exam Review Author: ghunter Last modified by: LCPS Created ... 237 Show detail 2 weeks ago quizzz.com Show details answer choices. a principal energy level. A circular path which electron travel around the nucleus. An area of high probability for finding an electron. a subatomic particle found outside the nucleus. Question 34. 900 seconds. Q. Calculate the average atomic mass of ... 215 Show detail 1 week ago lcps.org Show details Nuclear chemistry. Give the symbols for the following elements. a. B b. Be c. H d. He e. Beryllium. Boron. Helium. Hydrogen. Write the formula for the compound containing. a two to three ratio of iron to oxygen. a. Fe2O3 b. I2O3 c. Fe3O2 d. I3O2. six carbon, twelve hydrogen and six oxygen. a. CH2O b. C6H12O6 c. C6H6O6 420 Show detail 5 days ago lcps.org Show details Title: Chemistry Midterm Exam Review Author: ghunter Last modified by: LCPS Created Date: 1/8/2013 12:34:00 PM Company: LCPS Other titles: Chemistry Midterm Exam Review 185 Show detail 3 days ago quizzz.com Show details Play this game to review undefined. What is the variable the scientist wants to measure or observe? Chemistry Midterm Exam Review DRAFT. 9th grade. 0 times. 0% average accuracy. 5 hours ago. marketabrown. 0. Save. Edit. Edit. Chemistry Midterm Exam ... 306 Show detail 3 days ago quizzz.com Show details Play this game to review Chemistry. The group number tells us how many ____ there are in an element. The group number tells us how many ____ there are in an element. Chemistry Midterm exam review DRAFT. 10th - 12th grade. 0 times. Chemistry. 0% average accuracy. an hour ago. cry07281. 0. Save. Edit. Edit. Chemistry Midterm exam review ... 303 Show detail 1 week ago chemistry.coach Show details Solution. Question 6. Draw all possible products resulting from the substitution of one hydrogen of the compound F by one bromine atom via a radical process. Solution. 4 equivalent hydrogen groups = 4 products: Question 7. Follow the flow of electrons indicated by the curved arrows and predict the products of the reaction: 197 Show detail 1 day ago chemistry.coach Show details Midterm 1 includes 6 questions and it covers Chapters 1, 2 and 3 of General Chemistry 1. At the end of the test you can see the answers and determine whether you have solved the questions. You have 1 hour to resolve Midterm 1. 469 Show detail 3 days ago quizzz.com Show details Play this game to review Chemistry. What is the name for AlBr₃? Preview this quiz on Quizizz. What is the name for AlBr₃? Chemistry Midterm Exam Review DRAFT. 10th grade ... 0 times. Chemistry. 0% average accuracy. an hour ago. katiezebedeo. 0. Save. Edit. Edit. Chemistry Midterm Exam Review DRAFT. an hour ago. by katiezebedeo. Played 0 times ... 351 Show detail 4 days ago copley-fairlawn.org Show details AP Chemistry Midterm review packet. The test will be formatted like a mini AP Test. The exam starts at 7:00 so there 2 hours and 20 minutes to complete the exam. The period will be broken into 2 70 minute segments with a 5 minute break. One for multiple choice without a calculator, and one for free response where you are allowed to use a ... 421 Show detail 1 day ago coursehero.com Show details View Chemistry Midterm Exam ReviewMCI314.doc from SCIENCE 2003350 at Boca Raton Community High School. Chemistry Semester Exam Review Complete all problems on notebook paper. Transfer your final 480 Show detail 1 week ago coursehero.com Show details View Notes - chemistry-midterm-exam-review-sheet.jpg from CHEM 1211L at University Of Georgia. Chemistry Midterm Exam Review Sheet Spring 2012 Unit 1 - Scientific Method and Safety and Chapter 3 1. 227 Show detail 1 day ago quizzz.com Show details Glucose (C₆H₁₂O₆) Carbon dioxide (CO₂) Bromine (Br₂) Question 17. 30 seconds. Q. Chromium is characterized by a very high melting point. Use the Periodic Table to identify which of these elements most likely has a similar melting point. answer choices. 396 Show detail 1 week ago mylibrary.org Show details PDF Chemistry - Mid Term Exam Review Sheet #1...Review Sheet #1 The midterm exam covers chapters 1 - 4 & 9 - 11. You should read through each chapter, look over old tests you still have, answer the l-lowing Write a chemic al formula for each manle given and tell whether it is an ionic (i) or molecular (m) compound: a. 193 Show detail 3 days ago coursehero.com Show details View Notes - MidTerm Exam HChem Practice Exam Review with answers.doc from MATH 101 at Lambert High School. 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Chemistry - Mid Term Exam Review Sheet #1 The midterm exam covers chapters 1 - 4 & 9 - 11. You should read through each chapter, look over old tests you still have, answer the l-lowing questions and do the calculations in order prepare yourself lbr the mid-157 Show detail 1 week ago livingston.org Show details CP CHEMISTRY MIDTERM REVIEW PACKET ANSWER KEY. CHAPTER 2. a) State the law of conservation of matter and energy _____. In any chemical or physical process, matter and energy cannot be created or destroyed _____. b) Hydrogen reacts with oxygen to produce water. 2 H₂ + O₂ 2 H₂O. How many grams of hydrogen reacts with 32 grams of oxygen to ... 57 Show detail 1 day ago wtps.org Show details Honors Chemistry Midterm Exam Review. Topics: UNIT 1 - UNIT 5. Formula writing. Nomenclature. Molar mass and gram/mole conversions. Gram/mole/liter/molecule conversions. ... 10 point open-ended questions for you to answer on the midterm exam. The following are examples and practice to prepare for the midterm questions. The essay topics are ... 424 Show detail 1 week ago howard.edu Show details May 20, 2022 · Chemistry Midterm Exam Review Chemistry For Dummies, 2nd Edition (9781119293460) was previously published as Chemistry For Dummies, 2nd Edition (9781118007303). While this version features a new Dummies cover and design, the content is the same as the prior release and should not be considered a new or updated 275 Show detail 1 week ago kwanga.net Show details The honors chemistry midterm exam will consist of 125 multiple-choice questions for honors, covering the material from first semester. The exam is divided into sections by topics. ... This exam review was written directly from the exam. 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